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Example 1. How many molecules of SO_3 are needed to react with 144 molecules of Fe_2O_3 given this balanced chemical equation?. $\text{Fe}_2\text{O}_3 (s) + 3\text{SO}_3 (g) \rightarrow \text{Fe}_2(\text{SO}_4)_3$. Solution. We use the balanced chemical equation to construct

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a conversion factor between Fe_2O_3 and SO_3 . The number of molecules of Fe_2O_3 goes on the bottom of our conversion factor so it cancels with our given amount ...

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Practice Test Ch3 Stoichiometry (page 2 of 2) 19. The mass of element X found in 1.00 mole of each of four different compounds is 28.0 g, 42.0 g, 56.0 g, and 70 g, respectively. The possible atomic weight of X is a. 8.00 b. 14.0 c. 28.0 d. 38.0 e. 42.0 $N_2(g) + 2O_2(g) \rightarrow N_2O_4(g)$ 20. The above reaction takes place in a closed flask.

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Chapter Test B, continued 31. Calculate the energy released by freezing 13.8 g of a liquid. Its molar mass is 82.9 g/mol, and its molar enthalpy of fusion is 4.60 kJ/mol.

Assessment Chapter Test B - Ed W. Clark High School

CHAPTER 9 REVIEW Stoichiometry
SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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